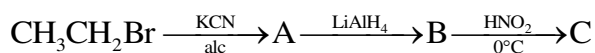


**KEY ANSWERS**

1	C	16	A	31	B	46	A
2	D	17	C	32	D	47	D
3	B	18	D	33	D	48	C
4	B	19	B	34	A	49	D
5	C	20	D	35	B	50	D
6	D	21	D	36	D	51	D
7	D	22	A	37	D	52	C
8	A	23	C	38	A	53	D
9	D	24	C	39	D	54	A
10	A	25	A	40	A	55	D
11	B	26	D	41	D	56	A
12	D	27	D	42	D	57	A
13	B	28	A	43	B	58	C
14	A	29	B	44	A	59	C
15	D	30	B	45	B	60	B

- In solid state,  $\text{PCl}_5$  is a/an  
 A) Octahedral structure  
 B) Ionic solid with  $[\text{PCl}_6]^+$  and  $[\text{PCl}_4]^-$   
 C) Ionic solid with  $[\text{PCl}_4]^+$  and  $[\text{PCl}_6]^-$   
 D) Covalent solid present in the form of  $\text{P}_2\text{Cl}_{10}$  **Ans. (C)**
- In which one of the following pairs, both the elements does not have  $(n-1)d^{10}ns^2$  configuration in its elementary state ?  
 A) Zn, Cd  
 B) Cd, Hg  
 C) Hg, Cu  
 D) Cu, Zn **Ans. (D)**
- Which of the following is CORRECT with respect to melting point of a transition element ?  
 A)  $\text{V} > \text{Cr}$   
 B)  $\text{Cr} > \text{Mn}$   
 C)  $\text{Mn} > \text{Fe}$   
 D)  $\text{Ti} > \text{V}$  **Ans. (B)**
- $a\text{MnO}_4^- + b\text{S}_2\text{O}_3^{2-} + \text{H}_2\text{O} \rightarrow x\text{MnO}_2 + y\text{SO}_4^{2-} + z\text{OH}^-$  a and y respectively are  
 A) 8; 3  
 B) 8; 6  
 C) 3; 6  
 D) 8; 8 **Ans. (B)**  
**Solution :**  $8\text{MnO}_4^- + 3\text{S}_2\text{O}_3^{2-} + \text{H}_2\text{O} \rightarrow 8\text{MnO}_2 + 6\text{SO}_4^{2-} + 2\text{OH}^-$
- Which formula and name combination is INCORRECT ?  
 A)  $\text{K}_3[\text{Al}(\text{C}_2\text{O}_4)_3]$  – Potassium trioxalatoaluminate (III)  
 B)  $[\text{Pt}(\text{NH}_3)_2\text{Cl}(\text{NO}_2)]$  – Diamminechloridonitrito – N – platinum (II)  
 C)  $[\text{CoCl}_2(\text{en})_2]\text{Cl}$  – Dichloridodiethylenediammine cobalt (II) chloride  
 D)  $[\text{Co}(\text{NH}_3)_4(\text{H}_2\text{O})\text{Cl}]\text{Cl}_2$  – Tetraammineaquachloridocobalt (III) chloride **Ans. (C)**
- Which of the following in an octahedral complex has maximum unpaired electrons ?  
 A)  $d^9$  (high spin)  
 B)  $d^6$  (low spin)  
 C)  $d^4$  (low spin)  
 D)  $d^7$  (high spin) **Ans. (D)**
- The correct decreasing order of basicity of hydrides of Group-15 elements is  
 A)  $\text{SbH}_3 > \text{AsH}_3 > \text{PH}_3 > \text{NH}_3$   
 B)  $\text{PH}_3 > \text{AsH}_3 > \text{SbH}_3 > \text{NH}_3$   
 C)  $\text{AsH}_3 > \text{SbH}_3 > \text{NH}_3 > \text{PH}_3$   
 D)  $\text{NH}_3 > \text{PH}_3 > \text{AsH}_3 > \text{SbH}_3$  **Ans. (D)**
- Which one of the following oxoacids of phosphorus can reduce  $\text{AgNO}_3$  to metallic silver ?  
 A)  $\text{H}_3\text{PO}_2$   
 B)  $\text{H}_4\text{P}_2\text{O}_7$   
 C)  $\text{H}_4\text{P}_2\text{O}_6$   
 D)  $\text{H}_3\text{PO}_4$  **Ans. (A)**  
**Solution :**  $\text{H}_3\text{PO}_2$  having 2P-H bonds
- A pair of compounds having the same boiling points are  
 A) cis but-2-ene and trans but-2-ene  
 B) n-hexane and neo-hexane  
 C) benzene and naphthalene  
 D) (+) butan – 2 – ol and (–) butan – 2-ol **Ans. (D)**
- Identify a, B and C in the sequence :



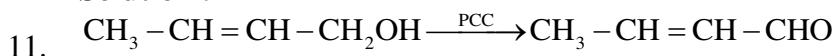
A)  $\text{CH}_3\text{CH}_2\text{CN}$ ,  $\text{CH}_3\text{CH}_2\text{CH}_2\text{NH}_2$ ,  $\text{CH}_3\text{CH}_2\text{CH}_2\text{OH}$

B)  $\text{CH}_3\text{CH}_2\text{NC}$ ,  $\text{CH}_3\text{CH}_2\text{OH}$ ,  $\text{CH}_3\text{CH}_2\text{CH}_2\text{NH}_2$

C)  $\text{CH}_3\text{CH}_2\text{CN}$ ,  $\text{C}_2\text{H}_5\text{OH}$ ,  $\text{C}_2\text{H}_5\text{N}_2\text{Cl}$

D)  $\text{CH}_3\text{CH}_2\text{CN}$ ,  $\text{CH}_3\text{CH}_2\text{NH}_2$ ,  $\text{C}_2\text{H}_5\text{OH}$

Ans. (A)



Hybridisation change involved at C-1 in the above reaction

A)  $\text{sp}^3$  to  $\text{sp}$

B)  $\text{sp}^3$  to  $\text{sp}^2$

C)  $\text{sp}^2$  to  $\text{sp}^3$

D)  $\text{sp}$  to  $\text{sp}^2$

Ans. (B)

12. If a didentate ligand ethane-1, 2-diamine is progressively added in the molar ratio

$\text{en} : \text{Ni} :: 1 : 1, 2 : 1, 3 : 1$  to  $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$  aq solution, following co-ordination entities are formed.

I.  $[\text{Ni}(\text{H}_2\text{O})_4\text{en}]^{2+}_{(\text{aq})}$  – pale blue

II.  $[\text{Ni}(\text{H}_2\text{O})_2(\text{en})_2]^{2+}_{(\text{aq})}$  – blue / purple

III.  $[\text{Ni}(\text{en})_3]^{2+}_{(\text{aq})}$  – violet

The wavelength in nm of light absorbed in case of I and III are respectively

A) 475 nm and 310 nm B) 300 nm and 475 nm C) 310 nm and 500 nm D) 600 nm and 535 nm

Ans. (D)

13. Which of the following is an organometallic compound ?

A)  $\text{CH}_3\text{COONa}$

B)  $\text{CH}_3\text{CH}_2\text{MgBr}$

C)  $(\text{CH}_3\text{COO})_2\text{Ca}$

D)  $\text{CH}_3\text{ONa}$

Ans. (B)

14. A better reagent to oxidize primary alcohols into aldehyde is :

A) PCC

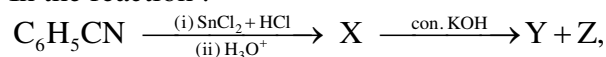
B) Alkaline  $\text{KMnO}_4$

C) Acidified  $\text{K}_2\text{Cr}_2\text{O}_7$

D)  $\text{CrO}_3$

Ans. (A)

15. In the reaction :



Formation of X, formation of Y and Z are known by

A) Rosenmund reduction, Cannizaro reaction

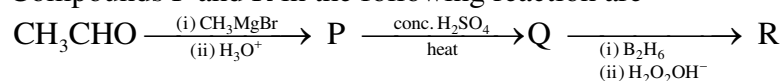
B) Clemmensen reduction, Sandmeyer reaction

C) Wolf-Kishner reduction, Wurtz reaction

D) Stephen reaction, Cannizaro reaction

Ans. (D)

16. Compounds P and R in the following reaction are



A) Position isomers

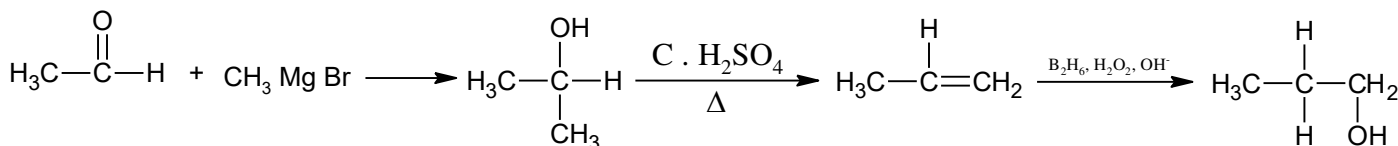
B) Functional isomers

C) Metamers

D) Identical

Ans. (A)

**Solution :**



17. Aniline does not undergo

A) Nitration

B) Sulphonation

C) Friedel-Craft reaction

D) Bromination

Ans. (C)

18. The heating of phenyl methyl ether with HI produces an aromatic compound A which on treatment with con.  $\text{HNO}_3$  gives B. A and B respectively are

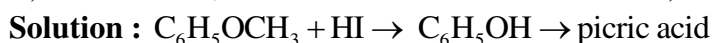
A) Methanol, Ethanoic acid

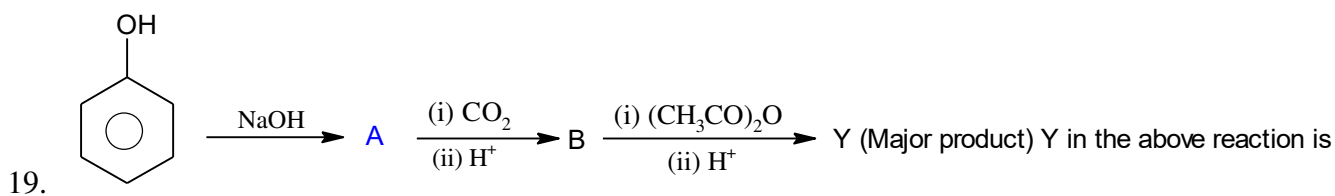
B) Peric acid, Phenol

C) Iodobenzene, 1-Iodo-4-nitrobenzene

D) Phenol, Picric acid

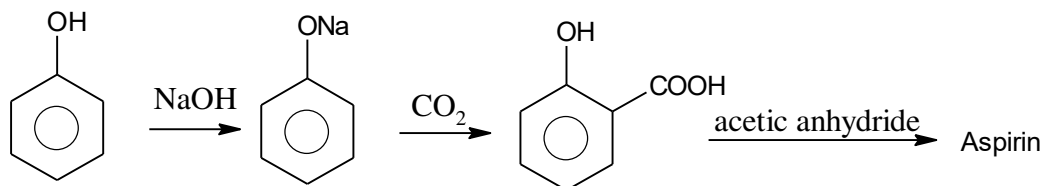
Ans. (D)





- A) Salicylaldehyde      B) Aspirin      C) Cumene      D) Picric acid      **Ans. (B)**

**Solution :**



20. Sucrose is dextrorotatory but after hydrolysis the mixture show laevorotation, this is because of

- A) Laevorotation of glucose is more than dextrorotation of fructose  
 B) Sucrose is a non-reducing sugar  
 C) Recemic mixture is formed  
 D) Laevorotation of fructose is more than dextrorotation of glucose

**Ans. (D)**

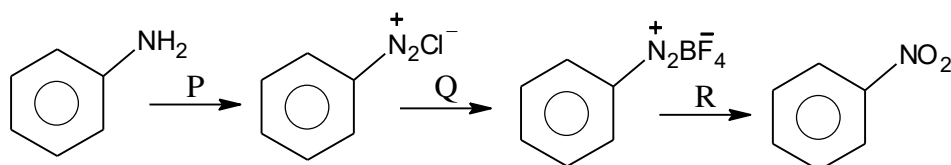
21. The correct order of match between column X and column Y is :

X	Y
I. Vitamin A	i. Muscular weakness
II. Vitamin D	ii. Increased blood clotting time
III. Vitamin E	iii. Night-blindness
IV. Vitamin K	iv. Osteomalacia

- A) I-iv, II-iii, III-ii, IV-i      B) I-ii, II-i, III-iii, IV-iv  
 C) I-iii, II-ii, III-iv, IV-i      D) I-iii, II-iv, III-i, IV-ii

**Ans. (D)**

22. In the reaction :



P, Q and R respectively are :

- A)  $\text{NaNO}_2 + \text{dil. HCl, HBF}_4, \text{Cu} + \text{NaNO}_2$       B)  $\text{NaNO}_2 + \text{con. HCl, F}_2, \text{Cu} + \text{NaNO}_3$   
 C)  $\text{NaNO}_2 + \text{dil. HCl, BF}_3, \text{Cu} + \text{NaNO}_2$       D)  $\text{NaNO}_3 + \text{dil. HCl, F}_2, \text{Cu} + \text{NaNO}_3$

**Ans. (A)**

23. Thyroxine produced in the thyroid gland is an iodinated derivative of \_\_\_\_\_

- A) threonine      B) lysine      C) tyrosine      D) tryptophan      **Ans. (C)**

24. Which one of the following is a non-narcotic analgesic ?

- A) Heroin      B) Codeine      C) Aspirin      D) Morphine      **Ans. (C)**

25. Receptors are proteins and crucial to body communication process. These receptors are embedded in

- A) Cell membrane      B) Protein      C) Endocrine gland      D) Chromosomes      **Ans. (A)**

26. Which of the following monomers form biodegradable polymers?

- A) Ethylene glycol and phthalic acid      B) Caprolactum and 1, 3-Butadiene  
 C) Phenol and formaldehyde      D) 3-hydroxybutanoic acid and 3-hydroxypentanoic acid

**Ans. (D)**

27. Match the list-I with list –II in the following :

List-I	List-II
1. Caprolactum	(a) $\begin{array}{c} \text{-(CH}_2\text{—CH)}_n \\   \\ \text{CH}_3 \end{array}$
2. Vinyl chloride	(b) $\begin{array}{c} \text{-(CH}_2\text{—CH)}_n \\   \\ \text{C}_6\text{H}_5 \end{array}$
3. Styrene	(c) $\begin{array}{c} \text{-(CH}_2\text{—CH)}_n \\   \\ \text{Cl} \end{array}$
4. Propene	(d) $\begin{array}{c} \text{O} \\    \\ \text{-(C—(CH}_2\text{))}_5\text{-N} \\   \\ \text{H} \end{array}$

A) 1-c, 2-d, 3-a, 4-b    B) 1-a, 2-d, 3-c, 4-b    C) 1-d, 2-c, 3-a, 4-b    D) 1-d, 2-c, 3-b, 4-a    **Ans. (D)**

28. The correct order of first ionisation enthalpy of given elements is

(A) Li < B < Be < C    (B) Be < Li < B < C    (C) C < B < Be < Li    (D) Li < Be < B < C    **Ans. (A)**

29. Which of the following statement is INCORRECT?

(A) Bond length of  $\text{O}_2 >$  Bond length of  $\text{O}_2^{2+}$     (B) Bond order of  $\text{O}_2^+ <$  Bond order of  $\text{O}_2^{2-}$   
 (C) Bond length of  $\text{O}_2 <$  Bond length of  $\text{O}_2^{2-}$     (D) Bond order of  $\text{O}_2 >$  Bond order of  $\text{O}_2^{2-}$     **Ans. (B)**

30. A gas at a pressure of 2 atm is heated from 25 °C to 323 °C and simultaneously compressed to  $\frac{2}{3}$  of its original value. Then the final pressure is

(A) 1.33 atm    (B) 6 atm    (C) 2 atm    (D) 4 atm    **Ans. (B)**

**Solution :**  $\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2} \Rightarrow \frac{2 \times V_1}{298} = \frac{P_2 \times \frac{2}{3} V_1}{596} \Rightarrow P_2 = 6$

31. Lattice enthalpy for NaCl is + 788 kJ mol<sup>-1</sup> and  $\Delta H_{\text{Hyd}}^\circ = - 784 \text{ kJ mol}^{-1}$ . Enthalpy of solution of NaCl is

(A) + 572 kJ mol<sup>-1</sup>    (B) + 4 kJ mol<sup>-1</sup>    (C) - 572 kJ mol<sup>-1</sup>    (D) - 4 kJ mol<sup>-1</sup>    **Ans. (B)**

**Solution :**  $\Delta H_{\text{soln}} = \text{L.E.} + \Delta H_{\text{Hydration}} = 788 - 784 = + 4 \text{ kJ}$

32. At 500 K, for a reversible reaction  $\text{A}_{2(\text{g})} + \text{B}_{2(\text{g})} \rightleftharpoons 2\text{AB}_{(\text{g})}$  in a closed container,  $K_C = 2 \times 10^{-5}$ . In the presence of catalyst, the equilibrium is attaining 10 times faster. The equilibrium constant  $K_C$  in the presence of catalyst at the same temperature is

(A)  $2 \times 10^{-4}$     (B)  $2 \times 10^{-6}$     (C)  $2 \times 10^{-10}$     (D)  $2 \times 10^{-5}$     **Ans. (D)**

33. A weak acid with  $\text{pK}_a$  5.9 and weak base with  $\text{pK}_b$  5.8 are mixed in equal proportions. pH of the resulting solution is

(A) 7.005    (B) 7.5    (C) 7    (D) 7.05    **Ans. (D)**

**Solution :**  $\text{PH} = 7 + \frac{1}{2}[\text{pka} - \text{pkb}]$

34. Temperature of 25, °C in Fahrenheit and Kelvin scale respectively are

(A) 77°F and 298.15K    (B) 17°F and 298.15K    (C) 45°F and 260.15K    (D) 47°F and 312.15 K    **Ans. (A)**

35. The number of protons, neutrons and electrons in the ion  $^{32}\text{S}^{2-}$  respectively are

(A) 16, 18, 16    (B) 16, 16, 18    (C) 18, 16, 16    (D) 16, 16, 16    **Ans. (B)**

36. A pair of amphoteric oxides is

(A)  $\text{Al}_2\text{O}_3, \text{Li}_2\text{O}$     (B)  $\text{BeO}, \text{BO}_3$     (C)  $\text{BeO}, \text{MgO}$     (D)  $\text{BeO}, \text{ZnO}$     **Ans. (D)**

37. The composition of water gas is

(A)  $\text{CO}_{(\text{g})} + \text{N}_{2(\text{g})}$     (B)  $\text{CH}_{4(\text{g})}$     (C)  $\text{CO}_{(\text{g})} + \text{H}_2\text{O}_{(\text{g})}$     (D)  $\text{CO}_{(\text{g})} + \text{H}_2_{(\text{g})}$     **Ans. (D)**

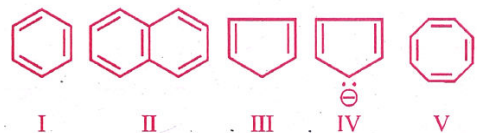
38. IUPAC name of the compound is



- (A) 2, 3 - dimethylbut-2-ene (B) 2, 3 - dimethyl butyne  
(C) 1, 1, 2, 2 - tetra methylethene (D) 2, 3 - dimethyl butane

Ans. (A)

39. Among the following:



The set which represents aromatic species is

- (A) I, II and III (B) III, IV and V (C) II and III (D) I, II and IV

Ans. (D)

40. Which one of the following gases converts haemoglobin into carboxy haemoglobin?

- (A) CO (B) O<sub>2</sub> (C) NO (D) CO<sub>2</sub>

Ans. (A)

41. What is the oxidation number of S in H<sub>2</sub>S<sub>2</sub>O<sub>8</sub>?

- (A) +5 (B) +4 (C) +7 (D) +6

Ans. (D)

**Solution :** Peroxodisulfuric acid

42. A 30% solution of hydrogen peroxide is

- (A) '30 volume' hydrogen peroxide (B) '10 volume' hydrogen peroxide  
(C) '50 volume' hydrogen peroxide (D) '100 volume' hydrogen peroxide

Ans. (D)

43. If 'a' stands for the edge length of the cubic systems - The ratio of radii in simple cubic, body centered cubic and face centered cubic unit cells is

- (A)  $1a : \sqrt{3}a : \sqrt{2}a$  (B)  $\frac{1}{2}a : \frac{\sqrt{3}}{4}a : \frac{1}{2\sqrt{2}}a$  (C)  $\frac{1}{2}a : \frac{\sqrt{3}}{2}a : \frac{\sqrt{2}}{2}a$  (D)  $\frac{1}{2}a : \sqrt{3}a : \frac{1}{\sqrt{2}}a$

Ans. (B)

44. Dimerisation of solute molecules in low dielectric constant solvent is due to :

- (A) Hydrogen bond (B) Covalent bond (C) Co-ordinate bond (D) Ionic bond

Ans. (A)

45. The swelling in feet and ankles of an aged person due to sitting continuously for long hours during travel, is reduced by soaking the feet in warm salt water. This is because of

- (A) Reverse Osmosis (B) Osmosis (C) Edema (D) Diffusion

Ans. (B)

46. A sample of water is found to contain 5.85%  $\left(\frac{w}{w}\right)$  of AB (molecular mass 58.5) and 9.50%  $\left(\frac{w}{w}\right)$  XY<sub>2</sub> (molecular mass 95). Assuming 80% ionisation of AB and 60% ionization of XY<sub>2</sub>, the freezing point of water sample is [Given : K<sub>f</sub> for water 1.86 K kg mol<sup>-1</sup>, Freezing point of pure water is 273 K and A, B and Y are monovalent ions]

- (A) 264.25 K (B) 265.56K (C) 280.44 K (D) 281.75 K

Ans. (A)

**Solution :**  $\Delta T = i k_f m$

$$\text{NaCl}, \alpha = \frac{i-1}{n-1}$$

$$\text{MgCl}_2, \alpha = \frac{i-1}{n-1}$$

$$0.8 = \frac{i-1}{2-1} = 1.8$$

$$0.6 = \frac{i-1}{3-1} = 2.2$$

$$\Delta T_f = iK_f m + iK_f m$$

$$= \frac{1.8 \times 1.86 \times 5.85 \times 1000}{58.5 \times 94.15} + \left( \frac{2.2 \times 1.86 \times 9.5 \times 1000}{90.5 \times 95} \right) = 8.071$$

$$\therefore T_f = 264.25 \text{ K}$$

47. Match the column A (type of crystalline solid) with the column B (example for each type):

A	B
P. Molecular Solid	i. Sic
Q. Ionic Solid	ii. Mg
R. Metallic Solid	iii. H <sub>2</sub> O
S. Network Solid	iv. MgO

(A) P-iii, Q-i, R-ii, S-iv

(B) P-iv, Q-iii, R-ii, S-i

(C) P-i Q-iv, R-iii, S-i

(D) P-iii, Q-iv, R-ii, S-I

**Ans. (D)**

48. A metal crystallises in a body centered cubic lattice with the metallic radius  $\sqrt{3} \text{ \AA}$ . The volume of the unit cell in  $\text{m}^3$  is

(A)  $64 \times 10^{-29}$

(B)  $4 \times 10^{-29}$

(C)  $6.4 \times 10^{-29}$

(D)  $4 \times 10^{-10}$

**Ans. (C)**

**Solution :**  $r = \frac{\sqrt{3}a}{4} \Rightarrow a = 4a, \text{ So } a^3 = 64 \times 10^{-30}$

49. The resistance of 0.1 M weak acid HA in a conductivity cell is  $2 \times 10^3 \text{ Ohm}$ . The cell constant of the cell is  $0.78 \text{ C m}^{-1}$  and  $\lambda_m^\circ$  of acid HA is  $390 \text{ S cm}^2 \text{ mol}^{-1}$ . The pH of the solution is

(A) 3.3

(B) 4.2

(C) 5

(D) 3

**Ans. (D)**

**Solution :**  $G^* = R \times K \Rightarrow 0.78 = 2 \times 10^3 \times K \Rightarrow K = 0.39 \times 10^{-3}$

$$\lambda = \frac{1000 \times K}{C} = 3.9$$

$$\alpha = \frac{\lambda m}{\lambda m^\circ} = \frac{3.9}{390} = 0.01$$

$$[\text{H}_3\text{O}^+] = C\alpha = 0.1 \times 0.01 = 10^{-3} \quad \therefore \text{pH} = 3$$

50. In which one of the following reactions, rate constant has the unit  $\text{mol L}^{-1} \text{ s}^{-1}$ ?

(A) Acid catalysed hydrolysis of  $\text{CH}_3\text{COOCH}_3$

(B)  $\text{CHCl}_3 + \text{Cl}_2 \rightarrow \text{CCl}_4 + \text{HCl}$

(C)  $2\text{NO}_{(\text{g})} + \text{O}_2 \rightarrow 2\text{NO}_{2(\text{g})}$

(D) Decomposition of HI on the surface of Gold

**Ans. (D)**

**Solution :** zero order reaction

51. For a reaction, the value of rate constant at 300 K is  $6.0 \times 10^5 \text{ s}^{-1}$ . The value of Arrhenius factor A at infinitely high temperature is

(A)  $6 \times 10^5 \times e^{-E_a/300}$

(B)  $e^{-E_a/300}$

(C)  $\frac{6 \times 10^5}{300}$

(D)  $6 \times 10^5$

**Ans. (D)**

52. The rate constants  $k_1$  and  $k_2$  for two different reactions are  $10^{16} \times e^{-2000/T}$  and  $10^{15} \times e^{-1000/T}$  respectively. The temperature at which  $k_1 = k_2$  is

(A)  $\frac{2000}{2.303} \text{ K}$

(B) 2000K

(C)  $\frac{1000}{2.303} \text{ K}$

(D) 1000K

**Ans. (C)**

**Solution :**

$$\log k_1 = 16 - \frac{2000}{2.303 T}$$

$$\log k_2 = 15 - \frac{1000}{2.303 T}$$

At  $k_1 = k_2$  option (C) is correct

53. During the electrolysis of brine, by using inert electrodes,

(A)  $\text{O}_2$  liberates at anode

(B)  $\text{H}_2$  liberates at anode

(C) Na deposits on cathode

(D) Cl<sub>2</sub> liberates at anode

Ans. (D)

54. Consider the following 4 electrodes

A : Ag<sup>+</sup> (0.0001 M)/Ag(s); B : Ag<sup>+</sup> (0.1 M) / Ag(s)

C : Ag<sup>+</sup> (0.01 M)/Ag(s); D : Ag<sup>+</sup> (0.00 1 M) / Ag(s);  $E^{\circ}_{Ag^+/Ag} = +0.80$  V

Then reduction potential in volts of the electrodes in the order

(A) B > C > D > A (B) C > D > A > B (C) A > D > C > B (D) A > B > C > D Ans. (A)

55. When FeCl<sub>3</sub> is added to excess of hot water gives a sol 'X'. When FeCl<sub>3</sub> is added to NaOH<sub>(aq)</sub> solution, gives sol 'Y'. X and Y formed in the above processes respectively are

(A) Fe<sub>2</sub>O<sub>3</sub> . xH<sub>2</sub>O / OH<sup>-</sup> and Fe<sub>2</sub>O<sub>3</sub> . xH<sub>2</sub>O/Fe<sup>3+</sup>

(B) Fe<sub>2</sub>O<sub>3</sub> . xH<sub>2</sub>O / H<sup>+</sup> and Fe<sub>2</sub>O<sub>3</sub> . xH<sub>2</sub>O/Na<sup>+</sup>

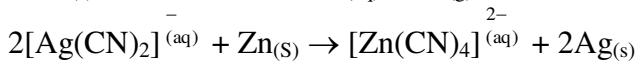
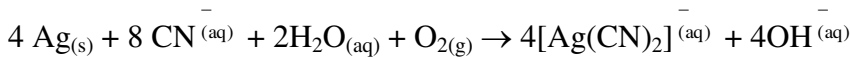
(C) Fe<sub>2</sub>O<sub>3</sub> . xH<sub>2</sub>O / Cl<sup>-</sup> and Fe<sub>2</sub>O<sub>3</sub> . xH<sub>2</sub>O/OH<sup>-</sup>

(D) Fe<sub>2</sub>O<sub>3</sub> . xH<sub>2</sub>O / Fe<sup>3+</sup> and Fe<sub>2</sub>O<sub>3</sub> . xH<sub>2</sub>O/OH<sup>-</sup>

Ans. (D)

**Solution :** A +vely charged sol of hydrated ferric oxide is formed due to adsorption of Fe<sup>3+</sup> ions

56. The reducing agent in the given equations:



(A) Zn

(B) O<sub>2</sub>

(C) H<sub>2</sub>O

(D) CN<sup>-</sup>

Ans. (A)

57. For the formation of which compound in Ellingham diagram ΔG° becomes more and more negative with increase in temperature?

(A) CO

(B) FeO

(C) ZnO

(D) Cu<sub>2</sub>O

Ans. (A)

58. Which of the following compound does not give dinitrogen on heating?

(A) Ba(N<sub>3</sub>)<sub>2</sub>

(B) NH<sub>4</sub>NO<sub>2</sub>

(C) NH<sub>4</sub>NO<sub>3</sub>

(D) (NH<sub>4</sub>) Cr<sub>2</sub>O<sub>7</sub>

Ans. (C)

**Solution :** NH<sub>4</sub>NO<sub>3</sub> → N<sub>2</sub>O + 2H<sub>2</sub>O

59. Aqueous solution of raw sugar when passed over beds of animal charcoal, it becomes colourless. Pick the correct set of terminologies that can be used for the above example.

	Adsorbent	Adsorbate	Process
(A)	Solution of Sugar	Animal Charcoal	Sorption
(B)	Animal Charcoal	Solution of Sugar	Absorption
(C)	Animal Charcoal	Colouring substance	Adsorption
(D)	Colouring Substance	Animal Charcoal	Adsorption

Ans. (C)

60. For Freundlich adsorption isotherm, a graph of log (x/m) Vs. log (P) gives a straight line. The slope of line and its Y-axis intercept respectively are

(A)  $\log \left( \frac{1}{n} \right), K$

(B)  $\frac{1}{n}, \log K$

(C)  $\log \left( \frac{1}{n} \right), \log K$

(D)  $\frac{1}{n}, K$

Ans. (B)

**Solution :**  $\log \frac{x}{m} = \log k + \frac{1}{n} \log p$